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1. (a) acid
- (b) acid
- (c) base
- (d) acid
- (e) base
- (f) acid
- (g) acid
- (h) acid
- (i) salt
- (j) base
- (k) base
- (l) salt
- (m) acid
- (n) salt
- (o) salt
- (p) salt
- (q) acid
- (r) acid
- (s) base
- (t) acid
- (u) acid
- (v) salt

Applying Knowledge**Acid-base neutralization reactions****Page 92**

1. (a) $\text{H}_2\text{SO}_4 + 2 \text{NaOH} \rightarrow 2 \text{H}_2\text{O} + \text{Na}_2\text{SO}_4$
- (b) $\text{HNO}_3 + \text{KOH} \rightarrow \text{H}_2\text{O} + \text{KNO}_3$
- (c) $2 \text{HCl} + \text{Ca}(\text{OH})_2 \rightarrow 2 \text{H}_2\text{O} + \text{CaCl}_2$
- (d) $2 \text{H}_3\text{PO}_4 + 3 \text{Ba}(\text{OH})_2 \rightarrow 6 \text{H}_2\text{O} + \text{Ba}_3(\text{PO}_4)_2$
- (e) $\text{CH}_3\text{COOH} + \text{NaOH} \rightarrow \text{H}_2\text{O} + \text{NaCH}_3\text{COO}$
- (f) $2 \text{HNO}_3 + \text{Sr}(\text{OH})_2 \rightarrow 2 \text{H}_2\text{O} + \text{Sr}(\text{NO}_3)_2$
- (g) $3 \text{HF} + \text{Fe}(\text{OH})_3 \rightarrow 3 \text{H}_2\text{O} + \text{FeF}_3$
- (h) $4 \text{HBr} + \text{Sn}(\text{OH})_4 \rightarrow 4 \text{H}_2\text{O} + \text{SnBr}_4$